Chemistry 115 Name

Dr. Cary Willard

Quiz 6a (20 points) October 22, 2013

1. (4 points)How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 4d orbital?
3. (2 points) How many electrons can occupy a 5p sublevel?
4. (4 points) What is the complete electron configuration of an atom of phosphorous?
5. (4 points) What is the shorthand electron configuration of an atom of technetium (Tc)?
6. (4 points) How does atomic size change as you move across the periodic table to the right? What is the explanation for this change?

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Quiz 6b (20 points) October 22, 2013

1. (4 points)How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 5p orbital?
3. (2 points) How many electrons can occupy a 4d sublevel?
4. (4 points) What is the complete electron configuration of an atom of silicon?
5. (4 points) What is the shorthand electron configuration of an atom of niobium (Nb)?
6. (4 points) How does atomic size change as you move across the periodic table to the right? What is the explanation for this change?

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Quiz 6c (20 points) October 22, 2013

1. (4 points)How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 5f orbital?
3. (2 points) How many electrons can occupy a 2p sublevel?
4. (4 points) What is the complete electron configuration of an atom of chlorine?
5. (4 points) What is the shorthand electron configuration of an atom of Yttrium (Y)?
6. (4 points) How does atomic size change as you move across the periodic table to the right? What is the explanation for this change?

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Quiz 6d (20 points) October 22, 2013

1. (4 points)How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 2p orbital?
3. (2 points) How many electrons can occupy a 5d sublevel?
4. (4 points) What is the complete electron configuration of an atom of sulfur?
5. (4 points) What is the shorthand electron configuration of an atom of cadmium (Cd)?
6. (4 points) How does atomic size change as you move across the periodic table to the right? What is the explanation for this change?